

### 4.5.3 Chemistry Paper 3 (233/3)

1 You are provided with:

- solution **A**, aqueous copper (II) sulphate;
- solid **B**, iron powder;
- 0.02 M acidified potassium manganate (VII), solution **C**.

You are required to determine the molar heat of displacement of copper by iron.

#### Procedure I

Using a burette, place 50.0 cm<sup>3</sup> of solution **A** in a 100 ml beaker. Measure the temperature of the solution and record it in table 1 below. Add **all** of solid **B** provided at once and start a stop watch. Stir the mixture **thoroughly** with the thermometer and record the temperature of the mixture after every one minute in the table. **Retain** the mixture for use in procedure **II** below.

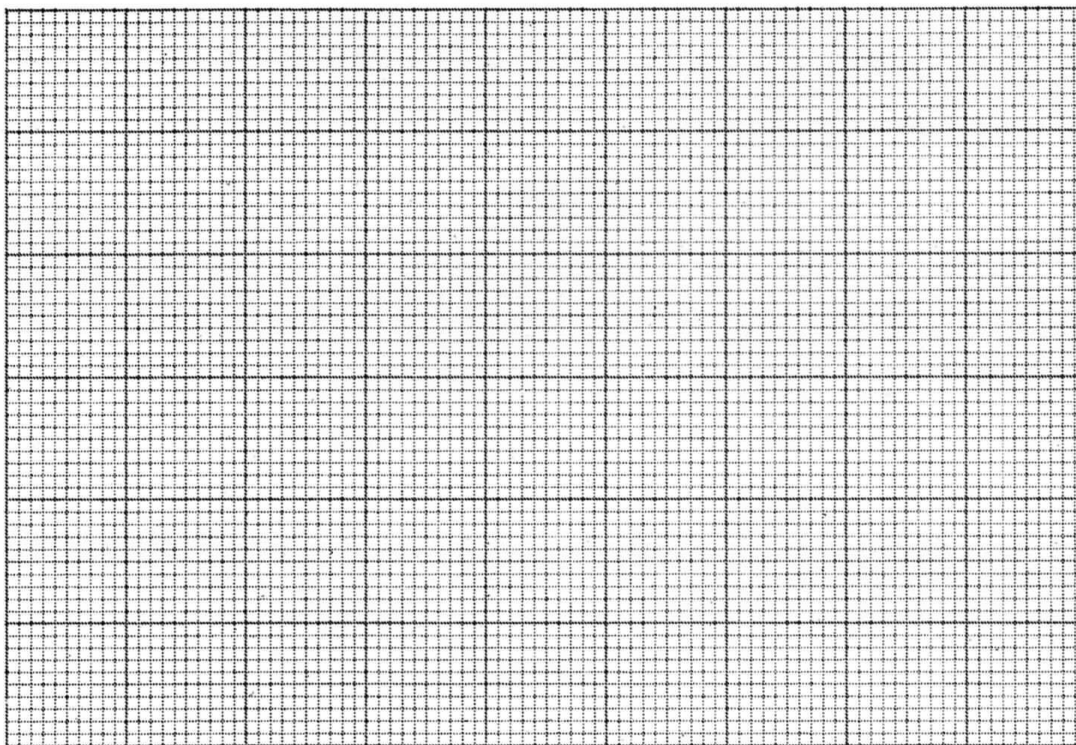
Table 1

Time (Min.)	0	1	2	3	4	5	6	7
Temperature (°C)								

(3 marks)

(a) (i) Plot a graph of temperature (vertical axis) against time in the grid provided.

(3 marks)



(ii) From the graph, determine the;

(I) highest change in temperature,  $\Delta T$ ;

(1 mark)

- (II) time taken for reaction to be completed. ( $\frac{1}{2}$  mark)
- (iii) Calculate the heat change for the reaction. (Specific heat capacity of solution is  $4.2 \text{ Jg}^{-1} \text{ K}^{-1}$ ; Density of the solution is  $1 \text{ gcm}^{-3}$ ). (2 marks)

### Procedure II

Carefully decant the mixture obtained in procedure I into a 250 ml volumetric flask. Add about  $10 \text{ cm}^3$  of distilled water to the residue in the 100 ml beaker. Shake well, allow the mixture to settle and carefully decant into the volumetric flask. **Immediately**, add about  $50 \text{ cm}^3$  of 2 M sulphuric (VI) acid to the mixture in the volumetric flask. Add more distilled water to make  $250.0 \text{ cm}^3$  of solution. Label this as solution **D**.

Fill a burette with solution **C**. Using a pipette and a **pipette filler**, place  $25.0 \text{ cm}^3$  of solution **D** into a 250 ml conical flask. Titrate solution **D** against solution **C** until the **first permanent pink** colour is obtained. Record your results in table 2 below. Repeat the titration two more times and complete the table. Retain the remaining solution **C** for use in question 3.

Table 2

	I	II	III
Final burette reading			
Initial burette reading			
Volume of solution <b>C</b> used ( $\text{cm}^3$ )			

(4 marks)

- (a) Determine the average volume of solution **C** used. ( $\frac{1}{2}$  mark)
- (b) Calculate the number of moles of:
- (i) aqueous potassium manganate (VII) used; (1 mark)
- (ii) iron (II) ions in  $25.0 \text{ cm}^3$  of solution **D**. (1 mole of  $\text{MnO}_4^-$  reacts with 5 moles of  $\text{Fe}^{2+}$ ). (1 mark)
- (iii) iron(II) ions in  $250 \text{ cm}^3$  of solution **D**. (1 mark)
- (c) Calculate the molar heat of displacement of copper by iron. (2 marks)

2 You are provided with solid **E**. Carry out the following tests and write your observations and inferences in the spaces provided.

- (a) Place **all** of solid **E** in a boiling tube. Add about  $10 \text{ cm}^3$  of distilled water and shake thoroughly. Filter the mixture into another boiling tube. **Retain** the filtrate for use in test 2(b) below. Dry the residue using pieces of filter papers.

- (i) Transfer about half of the dry residue into a dry test-tube. Heat the residue strongly and test any gas produced using a burning splint.

Observations	Inferences
(1 mark)	(1 mark)

- (ii) Place the rest of the residue in a dry test-tube. Add 4 cm<sup>3</sup> of 2M hydrochloric acid. **Retain** the mixture for test (iii) below.

Observations	Inferences
(1 mark)	(1 mark)

- (iii) To 2 cm<sup>3</sup> of the solution obtained in (ii) above, add 6 cm<sup>3</sup> of aqueous ammonia dropwise.

Observations	Inferences
(1 mark)	(1 mark)

- (b) (i) To 2 cm<sup>3</sup> of the filtrate obtained in (a) above, add about 3 cm<sup>3</sup> of aqueous ammonia (Excess).

Observations	Inferences
(1 mark)	(1 mark)

- (ii) To 2 cm<sup>3</sup> of the filtrate, add about 2 cm<sup>3</sup> of 2M hydrochloric acid.

Observations	Inferences
(1 mark)	(1 mark)

- (iii) To 2 cm<sup>3</sup> of the filtrate, add one or two drops of barium nitrate solution.

Observations	Inferences
(1 mark)	(1 mark)

3 You are provided with solid **G**. Carry out the tests in (a) and (b) and write your observations and inferences in the spaces provided. Describe the method used in part (c).

- (a) Place about one third of solid **G** on a **metallic** spatula and burn it in a Bunsen burner flame.

Observations	Inferences
(1 mark)	(1 mark)

- (b) Dissolve all of the remaining solid **G** in about 10 cm<sup>3</sup> of distilled water in a boiling tube. Use the solution for tests (b)(i), (ii) and (c).

- (i) Place 2 cm<sup>3</sup> of the solution in a test-tube and add 2 drops of acidified potassium manganate (VII); solution **C**.

Observations	Inferences
(1 mark)	(1 mark)

- (ii) To 2 cm<sup>3</sup> of the solution, add all of solid sodium hydrogen carbonate provided.

Observations	Inferences
(1 mark)	(1 mark)

- (c) Determine the pH of the solution obtained in (b) above.

Method used	Inferences
(2 marks)	(1 mark)