

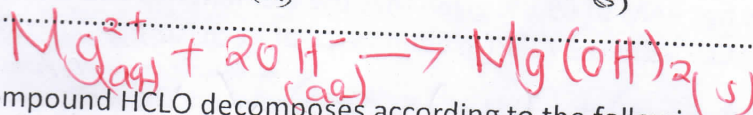
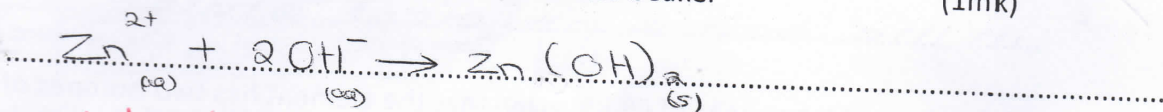
(b) State the observation in the beaker

(1mk) (2mk) (1mk)

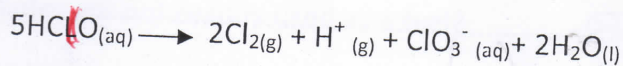
White ppt formed, which ~~durative~~ after some time

(c) Write ionic equation for the reaction in the beaker

(1mk)



19. The compound HClO decomposes according to the following equation.



Deduce the oxidation states of chlorine in the following species.

i) HClO

(1mk)

$$(+1) + \text{Cl} + (-2) = 0$$

$$\text{Cl} = +1$$

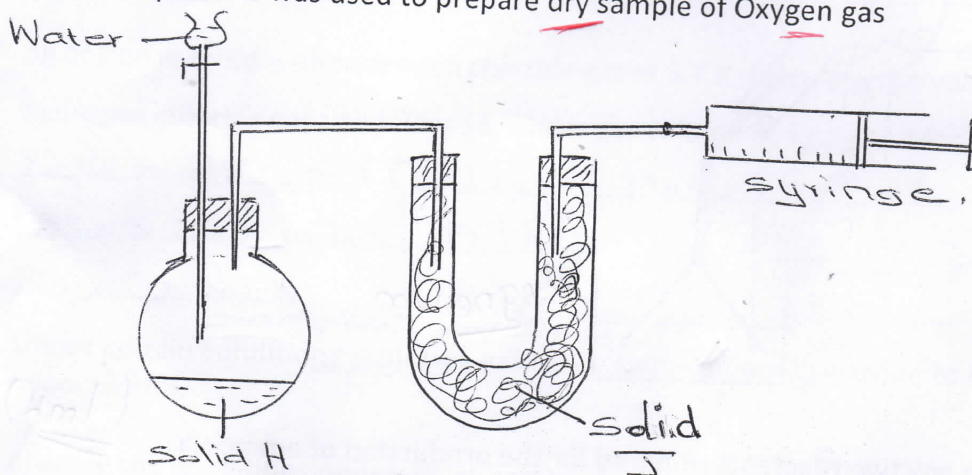
ii)  $\text{ClO}_3^{-}$

(1mk)

$$\text{Cl} + 3(-2) = -1$$

$$\text{Cl} = +5$$

20. The set-up below was used to prepare dry sample of Oxygen gas



(a) (i) Complete the diagram to show how the gas was collected

(1mk)

(ii) Identify the following

I. Solid H..... Sodium peroxide.....(1mk)

II. Solid J..... Anhydrous  $\text{CaCl}_2$ .....(1mk)