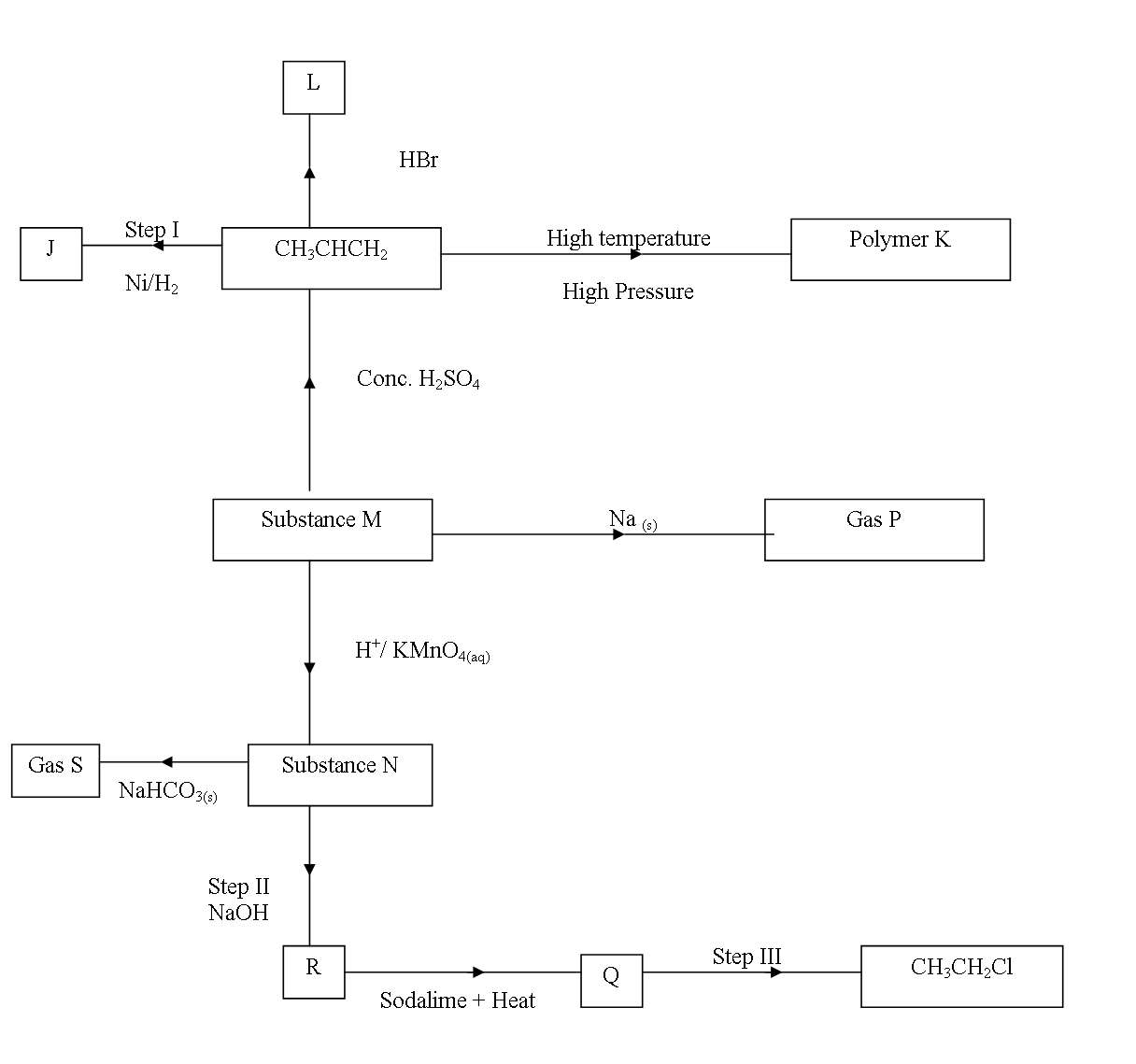
1. Use the flow chart below to answer the questions that follow.



(a) Name the following

(i) Gas S (1 mk)

(ii) Gas p (1 mk)

(iii) J (1 mk)

(b) Name process in

(i) Step I (1 mk)

(ii) Step II (1 mk)

(iii) Step III (1 mk)

(c) Draw two structural Isomers of compound L. (2 mks)

(d) Write a chemical equation for the complete combustion of Substance M. (1 mk)

(e) Name the reagent and condition in step III.

(i) Reagent. (1 mk)

(ii) Condition

(f) Calculate the mass of salt R that would be formed by using 21.9 tonnes of N when it reacts with excess Sodium hydroxide. (C=12.0, H=1.0, Na=23.0, O=16.0) (2 mks)

2 (b) The scheme below shows some reactions starting with Propanol. Study it and answer

the questions that follow.

S

CH3CH2COONa

CH3CH2CH2OH

B

CH3CH3

CH3CH=CH2

T

CH3CH2CH3

CH3CH2COOH

NaOH(aq)

Step II

K(S)

CH3COOH

Limited Cl2 and uv light.

Drops of Conc.H2SO4  Heat

Step I

Step III

Step IV

H H

C C

CH H n

i) Write down the formula of compounds S and T.

S\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_(1mk)

T\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_(1mk)

ii) Draw the structural formula of compound B. (2mks)

iii) Name the type of reaction, reagent and conditions in the reactions in step I and step IV.

|  |  |  |  |
| --- | --- | --- | --- |
| Step | Type of reaction | Reagent | Condition |
| I |  |  |  |
| IV |  |  |  |

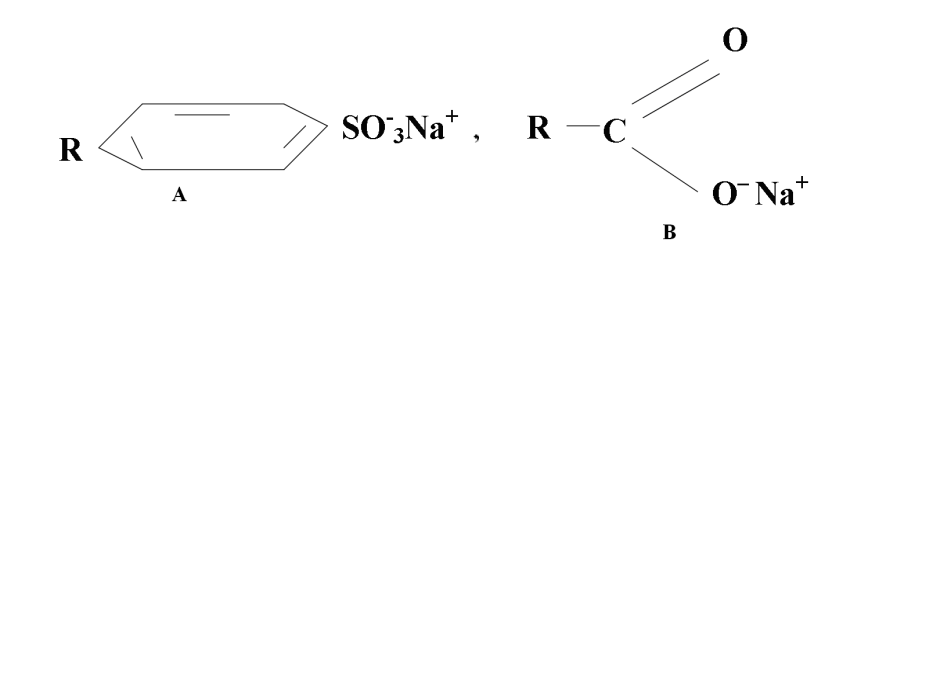
(3mks)

iv) ***Name two*** chemical tests that can be used to differentiate propanol from

propanoic acid.

|  |  |
| --- | --- |
| Propanol | Propanoic acid |
| i) |  |
| ii) |  |

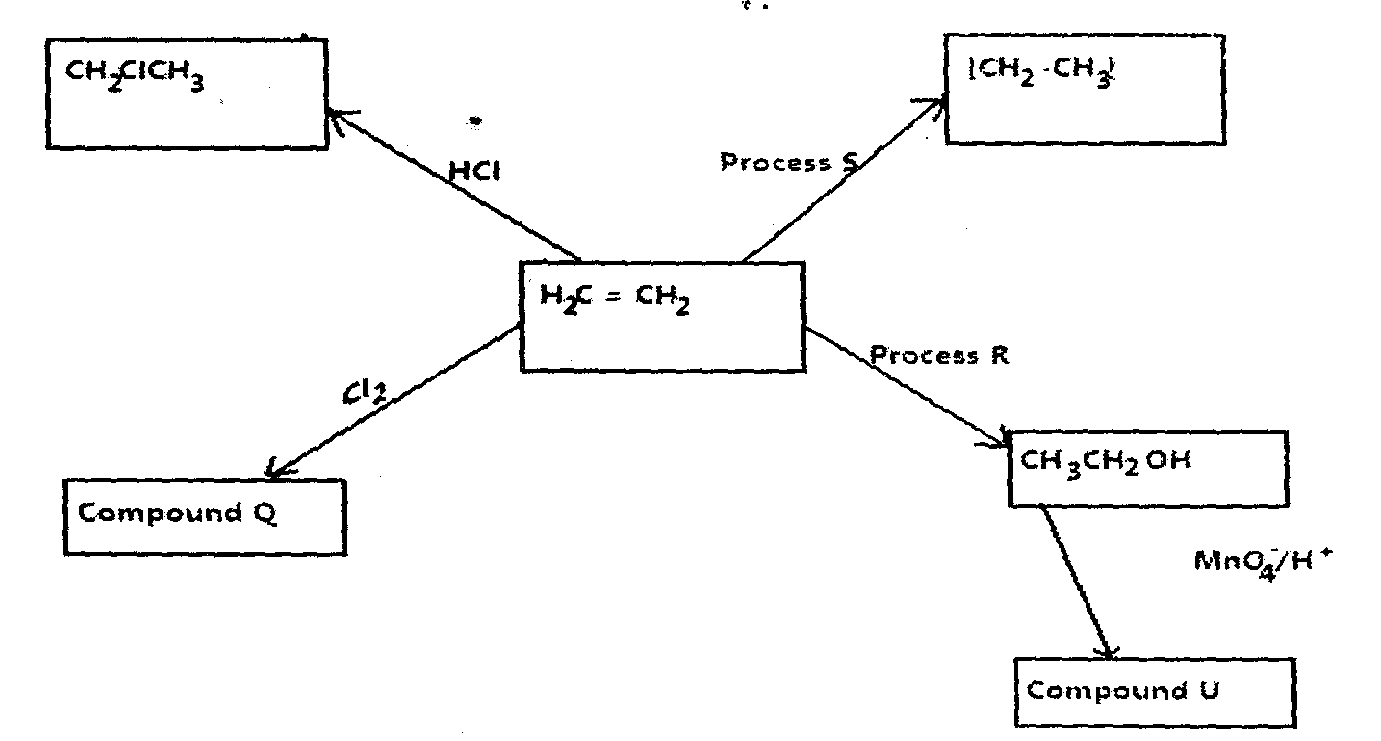
3. a) A and B shown below represent cleansing agents.



i) **State one** advantage of using A as a cleansing agent. (1mk)

ii) **Explain** how an aqueous solution of cleansing agent B removes oil from utensils during washing. (2mks)

b) Study the scheme below and answer the questions that follow.



i) **State** the:

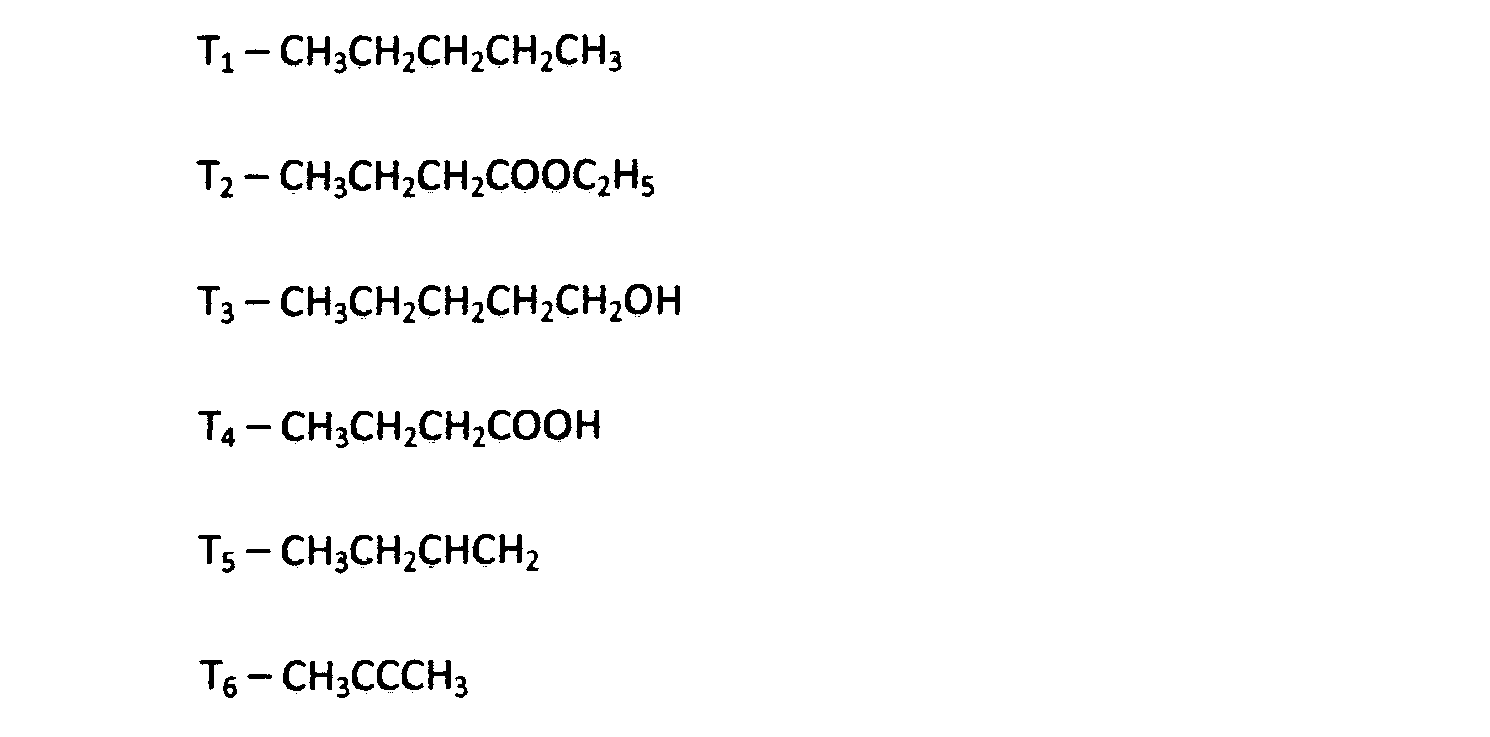
i) Conditions for process R

ii) Type of reaction representation by process S. (1mk)

iv) **Name** of compound U. (1mk)

ii) **Write** an equation for the formation of compound Q (1mk)

iii) **Draw and name** the structural formulae of compound CHClCH. (1mk)

4. The list below shows the formulae of some organic compounds. Use letters T1 to T6 to answer the questions that follow.   


a) Select two compounds which:

i) Are not hydrocarbons (lmk)

ii) Would decolourise both bromine water and acidified potassium manganate( VII). (lmk)

iii) Would produce hydrogen gas when reacted with potassium metal. (lmk)

b) Select a compound which would produce bubbles of a gas when reacted with sodium carbonate. (lmk)

c) i) Identify the compound that is likely to undergo polymerization. Give a reason

for your answer. Using two molecules show how polymerization occurs.   
I. Compound (lmk)

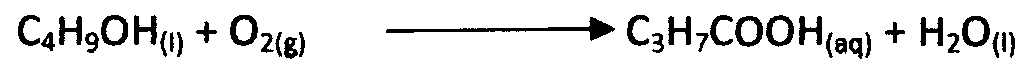
II. Reasons (lmk)

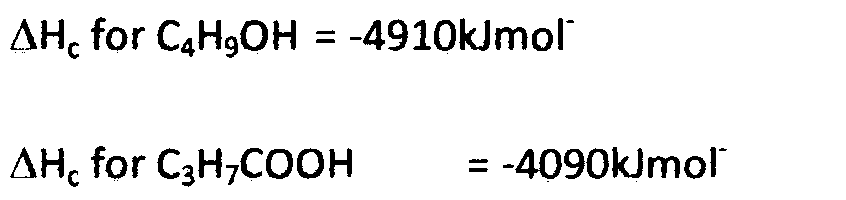
III. Polymerization (lmk)

ii) Name the process by which compound T2 is formed and identify the compounds   
that were used to form it.   
I. Process (lmk)

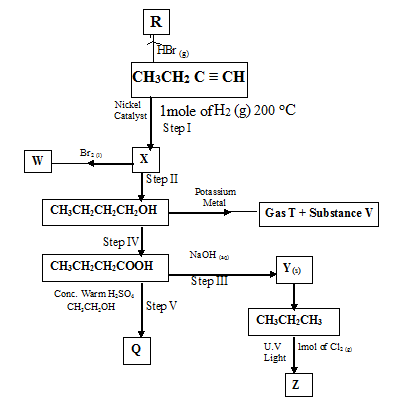
II. Compounds

i) ……………………………………………………………. (lmk)   
ii). ……………………………………………………………...

c) Compound T3 can be converted to T4 as shown by the equation below:   
 

Given the following information:   
  
Determine the heat change for the reaction above. (3mks)

5. The scheme below shows some reactions starting with but-1-yne. Study it and answer the questions that follow.



1. Name substances (2marks)

W\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

X\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Y\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

V\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

b) Name the type of reaction(s) taking place in; (3marks)

Step I\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Step III \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Step IV\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

1. State the reagents and conditions required for the formation of the following substances. (3marks)
2. CH3CH2CH2CH2OH

Reagent; \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Condition; \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

ii) CH3CH2CH3

Reagent; \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Condition; \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

iii) CH3CH2CH2COOH.

Reagent; \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Condition; \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

d) Write equations for the reactions that took place during the formation of;

i) R

1. W
2. Z

(3marks)

e) i) State a chemical test to differentiate between substance X and

CH3CH2CH3 (1mark)

ii) Ethanol (C2H5OH) dissolves readily in water yet dimethyl ether CH3OCH3, which has the same number and kind of atoms does not. Explain this observation. (2marks)

6.The scheme below shows several reactions starting with propanol. Study the scheme and answer the questions which follow.

**Propane**

**Q**

**CH2BrCHBrCH3**

**S**

**CH3CH2CH2OH**

**Gas R**

**CH3CH2 COOCH2 CH3**

**Bromine water**

**Excess**

**H+/K2Cr2O7**

**Na**

**Conc. H2SO4**

a) Name gas **R**. (1mk)

b) Name and draw the structural formula of compound **Q.** (2mks)

c) What conditions and reagents are necessary to convert **S** to CH3CH2COOCH2CH3 (2mks)

Reagents;

Conditions;

d) Write an equation for the reaction that takes place when equal volumes of chlorine gas react with propane. (1mk)

e) The table below shows some properties of organic compounds **U**, **V**, and **W**. use the information

to answer the questions that follow.

|  |  |  |  |
| --- | --- | --- | --- |
|  | **W** | **V** | **U** |
| Reaction with liquid bromine. | Decolourise bromine very fast | No reaction | Decolourises bromine liquid slowly |
| Combustion | Burns with yellow smoky flame | Burns with a blue flame leaving no residue | Burns with a yellow sooty flame |
| Reaction with conc. H2SO4 | No reaction | It is dehydrated to form compound **U**. | Reacts to form **V**. |

To which homologous series do the following compounds belong? (3mks)

**U**…………………………………………………………….

**V**…………………………………………………………….

**W**……………………………………………………………

f) CH2 CH – CH3 when heated under high temperatures and pressures forms a solid with large

molecular mass.

i) Write the equation for the reaction which involves the formation of the solid. (1mk)

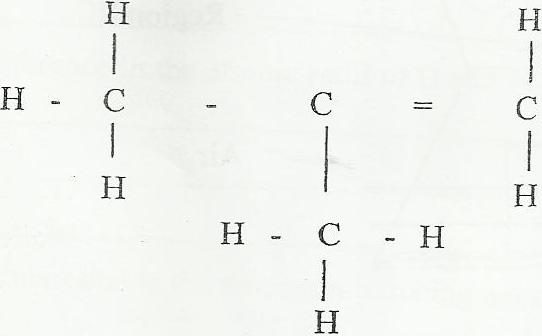
ii) Name the solid and give **one** use of the solid

Name (1mk)

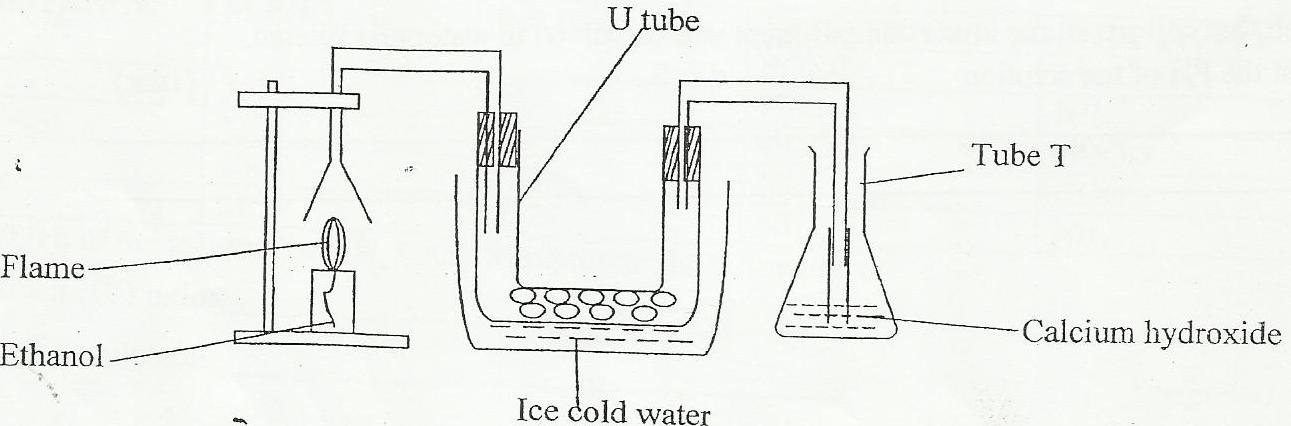
Use (1mk)

g) State **two** uses of cracking. (2mks)

1. (a) Give the IUPAC names of the following compounds. (1mk)

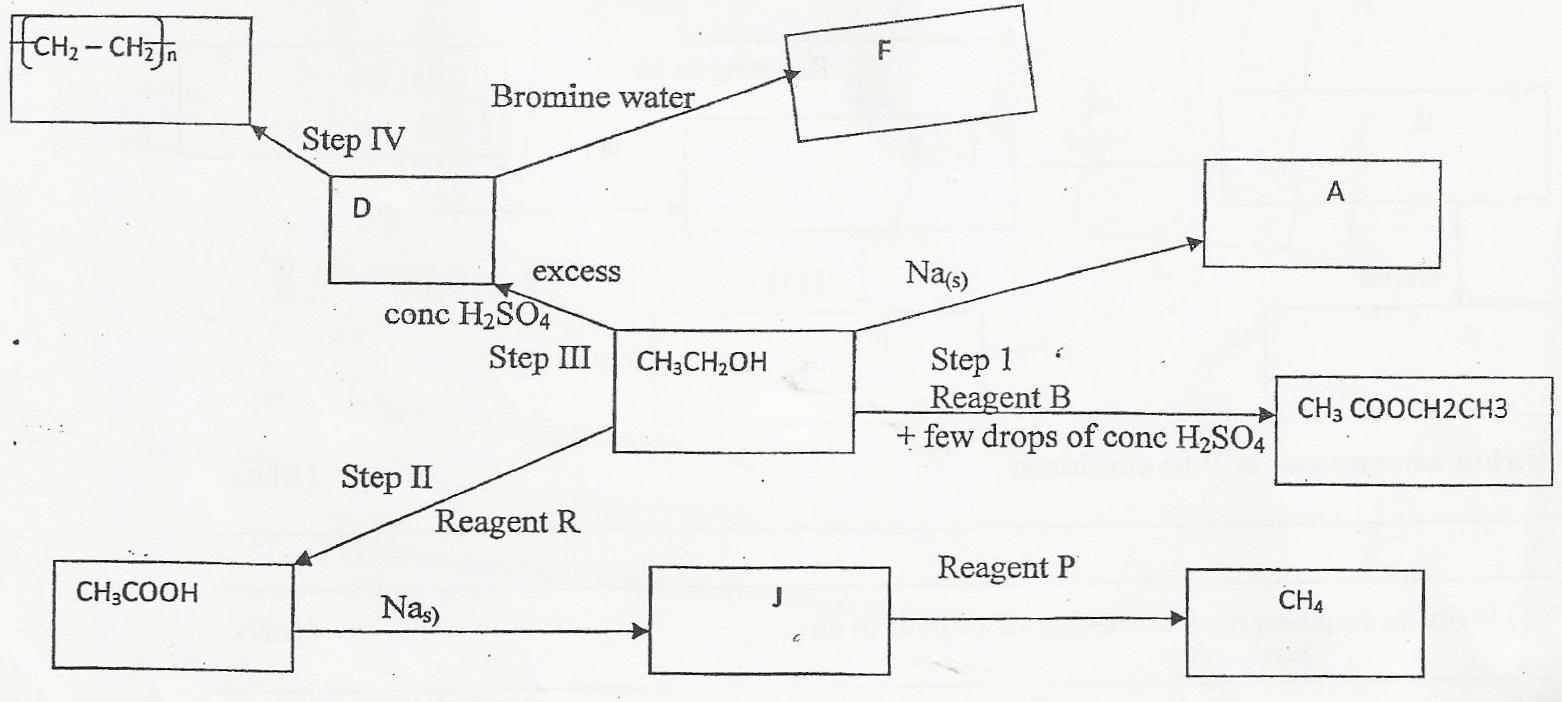
i)

ii) CH3 CH2COO CH2 CH3 (1mk)

(b) A lamp containing ethanol was burnt in air and the product collected as shown in the diagram.

State the observation made in the

1. U-tube (1mk)
2. Tube T (1mk)

(c) Study the reactions given in the diagram carefully and then answer the questions that follow.

1. Identify the following

I. Reagent B………………………………………………………… (1/2mk)

II. Reagent P…………………………………………………………. (1/2mk)

III. Reagent R ……………………………………………………… (1/2mk)

IV. Substance A…………………………………………………….. (1/2mk)

V. Substance F…………………………………………………….. (1/2mk)

VI. Substance D……………………………………………………….. (1/2mk)

ii. Write an equation to show formation of J. (1mk)

iii. Explain one disadvantage of the continued use of items made from compounds formed in step (IV)(1mk)

iv. What observation would be made when D is bubbled through bromine water? (1mk)

v. The compound D reacts with hydrogen in the presence of a catalyst G to form a compound H. Name:

I. Compound H…………………………………………………………………

II. Catalyst G……………………………………………………..……………….