**NAME………………………………………………...…………………………ADM NO.……………………………….**

**SCHOOL………………………………….…………………………………………………….…………..**

**STUDENT’S SIGN……………………....DATE………………………………………………………..**

**233/3**

**CHEMISTRY PRACTICAL**

**TERM TWO**

**CONFIDENTIAL**

**FORM THREE**

In addition to the apparatus and fittings found in the Chemistry Laboratory, each candidate will require the following;

1. 150cm 3 of solution C
2. 150cm3 of solution D
3. 60cm3 of Copper (II) sulphate (2 molar)
4. 50cm3Burette
5. 25cm3 pipette
6. Atleast 2 conical flasks
7. 100ml plastic beaker
8. (-100C = 1100C) thermometer
9. Spatula
10. One boiling tube
11. 5 test tubes
12. 1g of Zinc powder (accurately measured)
13. About 0.5g of solid Q
14. Tissue paper
15. Distilled water in a wash bottle

**Access to;**

* Phenolphthalein indicator
* Source of heat
* 2M sodium hydroxide with a dropper
* 2M ammonia solution with a dropper
* 2M hydrochloric acid with a dropper
* 2M Barium nitrate with a dropper
* 2M nitric (v) acid with a dropper
* 2M Lead (II) nitrate with a dropper

**NOTES**

1. Solution C is made by dissolving 10.08g/l of ethanedioic acid (oxalic acid) in 400cm3.
2. Solution D is made by dissolving 8.0g of Sodium hydroxide in 400cm3 of water and made to one litre.
3. Solid Q – aluminum sulphate (Al2(SO4)3)