NAME ………………………………………………………ADM NO. ……………………….

GATITU MIXED SECONDARY SCHOOL

CHEMISTRY FORM 2 TERM 3 2014 MID TERM EXAM

1. Study the flow chart below and answer the questions that follows



1. Name process 4mks

A

B

C

D

1. Give an example of a substance that undergo Nprocess A when heated. 1mk
2. Graphite is an allotrope of carbon. Explain why graphite is soft and slippery. 2mks
3. Sodium chloride do not conduct electricity in solid state but only in molten state. Explain.2mks
4. Define a salt and give two examples of salt. 2mks
5. Explain the following types of salts 4mks
6. Normal salt
7. Acidic salt
8. Double salt
9. Basic salt
10. Explain the following terms 3mks
11. Deliquence
12. Efflorescence
13. Hygroscopy
14. State two properties of the following

Ionic compounds 2mks

Covalent compounds 2mks

1. State four differences between luminous and non luminous flames. 4mks
2. Giving two reasons non –luminous flame is best to be used in laboratory heating.2mks
3. Most laboratory apparatus are made of glass. State two reasons why they are not made of plastic. 2mks
4. The grid below show a part of a periodic table. The letter does not represent the actual symbol. Use it to answer the questions that follow..



1. Compare the atomic atomic radius of element X and Y and give a reason for your answer. 2mks
2. Using crosses (x) to represent electrons draw the atomic structure of element Q 1mk
3. State the period and the group to which element Q belongs. 1mk
4. The ionic configuration of element A is 2.8 , A forms an ion of the type A- indicate on the grid the position of element A 1mk
5. During preparation of lead II chloride salt , lead II nitrate and sodium chloride were reacted in the laboratory using the equation below Pb(NO3)2 + NaCL → Pb cl2 + NaNo3
6. Name the insoluble salt that was formed 1mk
7. Balance the chemical reaction between lead II nitrate and sodium chloride. 2mks
8. Write an ionic equation to show how this salt was formed. 2mks
9. Name the type of reaction during formation of the above salt. 1mk
10. Name another salt that can be prepared by the similar method. 1mk
11. Which is the name given to ions that remain unchanged during a chemical reaction. 1mk
12. Identify the ions which didn’t take place in the above chemical reaction 1mk
13. The diagram below show a set up for preparation of oxygen gas.



1. Name solid M 1mk
2. Write a chemical reaction in the flask. 2mks
3. Why is oxygen collected by the above method? 2mks
4. Give two uses of oxygen gas. 2mks
5. Name a non metal that conduct electricity. 1mk
6. Explain why non metals do not conduct electricity. 2mks
7. Differentiate between conductors and electrolytes. 2mks