**FORM 4 MOCK CHEMISTRY PAPER 3**

**CONFIDENTIAL**

**Question one**

1. About 150cm3 solution H
2. About 100cm3 solution X
3. About 100cm3 solution N
4. One 50cm3 burette
5. One 25cm3 pipette
6. One filter funnel
7. Two 250ml conical flask

***NOTES***

1. Solution H is prepared by weighing accurately 3.16g of solid potassium manganite (VII), dissolve in about 400cm3 of 2M sulphuric (VI) acid and make the solution to 1 litre by adding distilled water. The solution should be prepared one day before the practical is taken.
2. Solution N is prepared by dissolving 24.5g of solid hydrated ammonium iron (II) sulphate in 200cm3 of 2M sulphuric (VI) acid and diluting the solution with distilled water to make 1 litre solution. This solution should be prepared in the morning of the examination.
3. Solution X is prepared by dissolving 5.0g of solid oxalic acid in 600cm3 of distilled water and making it up to one liter of solution with the distilled water.

**Question two**

1. About **80cm3** solution **B**.
2. About **120cm3** solution **C**
3. **250 m**l plastic beaker (empty)
4. **0 – 1100C** thermometer.
5. Stop watch

***NOTES***

1. Solution **B** is **1M** sodium hydroxide prepared by dissolving **40g** in **1** litre.
2. Solution **C** is **0.7M** sulphuric (vi) acid prepared by dissolving **38.5** litres of the acid in a litre of solution.

**Question three**

1.0.5M NaSO3 solution M

2. Oxalic acid R

**Access to**

1. Means of heating
2. Acidified potassium chromate (VI) supplied in a dropper
3. Distilled water
4. Litmus paper
5. NaOH
6. Bromine water
7. 0.5BaNO3
8. 1.0mHCl