CONFIDANTIAL **09**

1. $180cm^3$ of solution N HCl in a beaker.
2. $180cm^3$ of solution M 0.2M NaOH in a beaker.
3.. $50cm^3$ of solution D, 2M HCl in a beaker.
4. $100cm^3$ of solution B, 0.1M $Na_2S_20_3$ in a beaker.
5. $10cm^3$ of 1M HNO_3 in a boiling tube.
6. Ethanol in a stopped container.
7. $5cm^3$ of conc H_2SO_4 in a test tubewith a dropper.
8. Exact 1g of solid X which is F_2CO_3
9. About 1g of solid L in a dry stoppered container
10. About 1g of solid Q in a stoppered container access to:
11. Methy orange in a dropper
12. 0.5M $Ca(OH)_2$ in a dropper.
13. 1M $Pb(NO_3)_2$ in a dropper.
14. 2Ml NaOH solution in a dropper.
15. Distilled water in a wash bottle.
16. Acidified $KMnO_4$ in a dropper.
17. About 0.5g of Na_2CO_3 per student.
18. Pipette ($25cm^3$)
19. Burette.
20. Pipette filler.
21. 3 conical flasks.($250cm^3$)
22. Stand and clamp.
23. White tile.
24. $100cm^3$ glass beaker.
25. Thermometer (-10 to 110^0C)
26. $10cm^3$ measuring cylinder.
27. $100cm^3$ measuring cylinder.
28. Stop watch/ clock.
29. Plain white paper.
30. 2 boiling tube.
31. 6 test tubes in a test tube rack.
32. A glass rod.
33. Metallic spatula.
34. source of heat.

N/B

1. solid X, Solid L and solid Q to be provided by the examining authority.

Preparation of solutions

Solution N is prepared by dissolving 68.8ml of concentrated Hydrochloric acid in 500ml of distilled then top up to 1litre of solution.

2. 1M HNO_3 is prepared by 66ml of conc. HNO_3 in 500ml of distilled water then top up to make 1litre of solution

3. Acidified $KMNO_4$ is prepared by dissolving 3.16g of $KMnO_4$ in $400cm^3$ of 2M H_2SO_4 then topped up to 1litre of solution by distilled H_2O.

4. 2M H_2SO_4 prepared by dissolving $110cm^3$ of conc. H_2SO_4 in 500ml of distilled H_2O then top to 1litre of solution.