NAME…………………………………………………………………………………………….ADM. NO…………………….CLASS………….

**ST CLAIRE GIRLS’ HIGH SCHOOL ,GATITU.**

**P.O. BOX 327-01030 GATUNDU.**

**CHEMISTRY FORM TWO**

**TUNE-UP EXAMINATION TERM TWO 2017**

 INSTRUCTIONS

Answer all the questions in this paper.

 QUESTIONS.

1.a)Write the chemical symbols of the following elements:

i)Cobalt

ii)Copper

iii)Phosphorous

iv)Pottasium

v)Silicon (5mks)

b)Given the elements with the atomic number below .Write their electron configuration.(Letters do not represent the actual symbols of the elements)

L (7)

M (10)

K (16)

N (18)

 (4mks)

2.a)Define the following terms as used in Chemistry:

i)Ion

ii)Ionisation

iii)Cation

iv)Anion

 (4mks)

3.Below are the atomic number of five different elements .Note that the symbols are not the actual symbols of the elements .

|  |  |
| --- | --- |
| Element | Atomic number |
| A | 9 |
| B | 20 |
| C | 19 |
| D | 12 |
| E | 11 |

a)Identify an element which is a non-metal . (1mk)

b)Identify any two elements that are in the same group and state the group they belong.(2mks)

c)What is the name given to group which B belong . (1mk)

4.Give the valencies of cation and anion in each of the following compounds.

i)Zinc sulphate

ii)Magnesium carbonate

iii)Alluminium nitrate

 iv)Pottasium chloride

 (4mks)

5.An isotope Q has 18 neutrons and mass of 34 .

a.i)Write the electron arrangement of Q. (1mk)

ii)Draw the atomic structure of Q. (2mks)

b)To which period and group does Q belong ?Explain your answer . (2mks)

c)How does Q form its ion? Explain your answer . (1mk)

6.Determine the relative atomic mass of an element whose isotopic composition occurs in the proportions given.

 36 Ar(0.34%), 38 Ar(0.06%) and 40 Ar(99.6%) (3mk)