1.A compound has 37.5% carbon ,12.5% Hydrogen and the rest is oxygen.If its relative formular mass is twice its empirical formula mass,calculate its molecular formula and the relative molecular mass.

2.21g of a compound contains 8.4g carbon,1.4g Hydrogen and the rest is oxygen.What is its molecular formular if its relative molecular mass is 180?

3.Find the empirical formula of an oxide of metal X(r.a.m 56)containing 72.42% X and 27.58% oxygen

4.0.045 mole of a certain hydrocarbon gave on complete combustion 9.9g of carbon (iv)oxide and 4.86g of water.Calculate its molecular formula.

5.When a certain hydrocarbon burnt completely in excess oxygen,5.28g of carbon(iv)oxide and 2.16g of water were formed .If the molecular mass of the hydrocarbon is 84,determine the molecular formula of the hydrocarbon.