**Form IV Chemistry Practical Confidential term 2 2019**

1. Two 250ml conical flasks

2. One pipette filler

3. Two labels

4. 3 filter papers

5. About 75cm3 of solution **W9**

6. About 100cm3 of solution **W12**

**7.** About 150cm3 of solution **W11 (oxalic acid)**

**8.** Exactly 10cm of metal **M (magnesium ribbon)**

9. One 50cm3 burette

10. One 25cm3 pipette

11. One filter funnel

12. Thermometer

13. 0.5g grams of solid F

14. 0.5g of solid E

15. Aluminium foil

16. Red and blue litmus papers

17. 15cm3 of 2M HCl

18. About 0.3g of NaHCO3

19. 200cm3 of distilled water

20. Six dry test-tubes in a rack

21. Two boiling tubes

22. Tissue paper

23. Test-tube holder

24. Metallic spatula

Access to:

 Dilute sulphuric (V) acid

 2M NaOH

 Bromine water

 Bunsen burner

 Phenolphthalein indicator

**Preparations**

i. Solution W9 is made by dissolving 90cm3 of concentrated hydrochloric acid in distilled water and making it to one litre of solution. This solution MUST be supplied in a burette placed at a central position where it should be accessible to 5 to 10 candidates.

ii. Solution W11 is made by dissolving 6.3g of solid oxalic acid in distilled water and making it up to one litre of solution.

iii. Solution W12 is made by dissolving 3.2g of sodium hydroxide pellets in distilled water and making it up to one litre of solution.

iv. Metal M is magnesium ribbon should be cleaned with sand-paper the day before the examination.

v. Solid F is sodium benzoate

vi. Solid E is magnesium nitrate