CHEMISTRY CONFIDENTIAL

Each student requires

- 80 cm^3 of solution A prepared by dissolving 4.0g of NaOH in water made up to 1 litre.
- 200cm^3 of solution B which is 0.1M HCl.
- 2.5 of a salt mixture prepared by mixing 1.5 g of sodium carbonate (Anhydrous) and 1.0g of sodium chloride.
- One burette (50ml)
- One 25cm^3 pipette
- Pipette filler
- Complete stand
- Filter funnel
- White tile
- 3 conical flasks
- 250ml volumetric flask.
- Labels (6)
- 500ml distilled water in a wash bottle.
- 200cm^3 of solution S_3 which is sodium thiosulphate of concentration 0.2 M
- 80cm of 2m hydrochloric acid solution
- 100ml empty glass beaker.
- One stop watch/ clock.
- white piece paper
- One 50ml measuring cylinder
- About 2g of solid D (A mixture of ammonium sulphate and zinc sulphate in the ratio 1:1).
- About 5cm^3 of liquid B which is absolute ethanol
- Two red litmus paper.
- One metallic spatula.
- About 1g of solid sodium hydrogen carbonate
- Test- tube holder.
- One boiling tube
- Seven clean dry test tubes.

Provide access to:

- Means of heating
- Phenolphthalein indicator
- Methyl - indicator
- Acidified potassium dichromate (vi) solution.
- 2m nitric (V) acid solution
- 2m Barium nitrate solution
- 2m ammonia solution.
- 2m sodium hydroxide solution.
- 2m lead nitrate solution

Note: The solution should be supplied with droppers.